

## Chapter 9 Review Stoichiometry Answers Section 1

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### Chapter 9 Review Stoichiometry Answers

CHAPTER 9 REVIEW Stoichiometry MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation:  $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$  a. What is the value of the coefficient x in this equation? 40.07 g/mol b. What is the molar mass of  $C_3H_4$ ? 2 mol O<sub>2</sub>:1 mol H<sub>2</sub>O c. What is the mole ratio of O<sub>2</sub> to H

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### CHAPTER 9 REVIEW Stoichiometry

CHAPTER 9 REVIEW. Stoichiometry. MIXED REVIEW. SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation:  $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$  a. What is the value of the coefficient . x. in this equation? b. What is the molar mass of  $C_3H_4$ ? c. How many moles are in an 8.0 g sample of  $C_3H_4$ ? 2. a. What is meant by . ideal conditions

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### **Chemistry 9th Edition Chapter 3 - Stoichiometry - Review ...**

CHAPTER 9 REVIEW. Stoichiometry. SECTION 9.2. PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas:

### **CHAPTER 9 REVIEW**

fewer steps are required to solve stoichiometry problems when. ... Chemistry Chapter 9 Stoichiometry Test Review. 38 terms. Valerie\_a\_Chem CH 10. 55 terms. megfre186. Chemistry Chapter 6: Chemical Bonding. 30 terms. bluetejal12. Chemistry Chapter 4 Test. 50 terms. Briana\_Hanlon. Subjects. Arts and Humanities.

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Stoichiometry b. Theoretically, how many moles of  $\text{NH}_3$  will be produced? PROBLEMS Write the answer on the line to the left, Show all your work in the space provided. 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of  $\text{N}_2$  are mixed with 12.0 mol of  $\text{H}_2$  according to the ...

### **Date. FCHAPJ REV[EW.**

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### **Modern Chemistry Chapter 9 Mixed Review Stoichiometry Answers**

Composition stoichiometry deals with the mass relationships of elements in compounds. Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction. Reaction stoichiometry, the subject of this chapter, is based on chemical equations and the law of conservation of mass. All reaction stoichiometry

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reactant. Define product. Identify the products and reactants in a reaction. Identify a chemical change. Relate the symbols in a chemical equation to the words in a word equation. Write the word equation from a ...

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CHAPTER 9 REVIEW Stoichiometry SECTION 2 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 4.5 mol The following equation represents a laboratory preparation for oxygen gas:  $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$  How many moles of  $\text{O}_2$  form if 3.0 mol of  $\text{KClO}_3$  are totally consumed?

### **Modern Chemistry Chapter 9 Review Stoichiometry Section 2 ...**

CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left Show all your work in the space provided 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the percentage yield 2 60 mol of  $\text{N}_2$  are mixed with 120 mol of  $\text{H}_2$  according to the following equation:  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g})$

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