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Chapter 10 Chemical Quantities Practice Problems Answers

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Use the chemical equation to determine how many moles of water are present. Total the mass for all elements that make up one mole of water. Determine how many grams are present in 2 moles of water. Convert 36.04 g of water to kilograms using the conversion method.: e y . . The coefficient in front of H_2O is 2, so 2 moles of water are present.

Chemical Quantities - AP Biology

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CHAPTER 10: Chemical Quantities BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to 6.02×10^{23} particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

CHAPTER 10: Chemical Quantities

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92 Guided Reading and Study Workbook CHAPTER 10, Chemical Quantities(continued) 9. Complete the table about representative particles and moles. The Mass of a Mole of an Element (pages 293–294) 10. What is the atomic mass of an element? 11. Circle the letter of the phrase that completes this sentence correctly. The atomic masses of all elements

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287-296)

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