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## **Chapter 13**

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States of Matter.

STUDY. PLAY. kinetic

energy. the energy an

object has because of

its motion. kinetic

theory. a theory

explaining the states of

matter, based on the

concept that all matter

consists of tiny

particles that are in

constant motion.

kinetic theory for gases

**Chemistry Chapter**

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There are three states of matter that we will learn about in this chapter. (If you want to learn about more states of matter, I can refer you to somebody.) Those three states are solid, liquid, and gas. These three states are quite different. The main difference is in their particles.

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Section 13.4 Changes of State

- The relationship among the solid, liquid, and vapor states (or phases) of a substance in a sealed container are best represented in a single graph called a phase diagram
- Phase diagram - gives the temperature and pressure at which a substances exists as

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solid, liquid, or gas  
(vapor) </li> </ul>

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Chapter 13“States of  
Matter” 2. Section 13.1  
The Nature of Gases□  
OBJECTIVES: □Describe  
the assumptions of the  
“kinetic theory” as it  
applies to gases. 3.

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chemistry. in a system at constant vapor pressure, a dynamic equilibrium exists between the vapor and the liquid. The system is in equilibrium because the rate of evaporation of liquid equals the rate of condensation of vapor.

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States Of Matter. All matter consists of tiny particles that are in constant motion. Used to explain the properties of matter in terms of the energy and movement of the particles and the forces acting between them.

- 1.

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ISBN-10: 0132525763,

ISBN-13:

978-0-13252-576-3,

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- Page 444: 58 Answer

It is called a dynamic equilibrium because vaporization and condensation are still occurring at equal rates even though there is no net change in the number of particles in the liquid or vapor.

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Chapter 13 - States of Matter. Chapter 13 contains the following units: Gases. Forces of Attraction. Liquids and Solids, Phases

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**Chapter 13 - States  
of Matter - Ms. K  
Kelly - John F ...**

Chapter 13: gases For a fixed amount of gas, a change in one variable, pressure, temperature, or volume, affects the other two. The ideal gas law relates the number of particles to pressure, temperature, and volume. When gases react, the coefficients

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in the balanced chemical equation represent both molar amounts and relative volumes.

## **Chapters 12 & 13: States of Matter and Gases - ANNE ...**

Chapter 23 - Transition Metals and Coordination Chemistry - pp. 64-91; Chapter 23 - Transition Metals and Coordination Chemistry - pp. 92-114. If you're using the 10th or 11th



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editions this information is in 23.7 and ch 24. The homework links for those editions reflect this. Various oxidation states of Vanadium.  
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